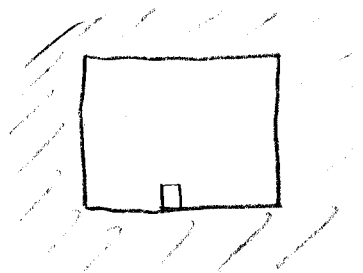


Esercizio del bicchiere di vapore acqua in una stanza

1



Contenitore con 10^2 g di H_2O allo stato gassoso. Contenitore inizialmente chiuso. Stanza a

Si stappa il contenitore. Il vapore condensa e si raffredda. Trovare la T_f del sistema contenitore + stanza

$$\left\{ \begin{array}{l} P_0 = 10^5 \text{ Pa} \\ V_0 = 3 \times 3 \times 2 \text{ m}^3 \\ T_0 = 293 \text{ K} \end{array} \right.$$

1) $n_0 = \frac{P_0 V_0}{R T_0}$ # moli di N_2 nella stanza. ^{supponiamo}

$m_{H_2O} = m_0$ quanti moli sono?

$$n = \frac{m(g)}{P_m} \quad P_m \stackrel{\text{def}}{=} \frac{\text{massa molecolare}}{\frac{1}{12} m(^{12}C)}$$

$$m(^{12}C) = 19.93 \times 10^{-24} \text{ g} \Rightarrow \frac{1}{12} m(^{12}C) = 1.66 \times 10^{-24} \text{ g}$$

$$H_2O = 2H + O \Rightarrow P_m(H_2O) = 2P_m(H) + P_m(O) \approx 18$$

$$\Rightarrow n = \frac{10^2}{18} \approx 5.56$$

Nota $n_0 \approx 740$!!

2) $Q_{H_2O \rightarrow \text{stanza}} + Q_{\text{stanza} \rightarrow H_2O} = 0$

$$Q_{H_2O \rightarrow st} = -m_{H_2O} \lambda_e + m_{H_2O} C_{H_2O} (T_f - T_e)$$

$$Q_{st \rightarrow H_2O} = \underbrace{m_0 C_p}_{m_0 c_p} (T_f - T_0)$$

$$\Rightarrow$$

$$T_f = \frac{m_o C_p T_o + m_{H_2O} c_{H_2O} T_e + m_{H_2O} \lambda_e}{m_{H_2O} c_{H_2O} + m_o C_p}$$

È interessante mettere i numeri

$$m_o C_p T_o \approx 740 \times \frac{7}{2} \times 2 \frac{\text{cal}}{\text{K}} \times 293 \text{ K} = 1.52 \times 10^5 \text{ cal}$$

$$m_{H_2O} c_{H_2O} T_e \approx 10^2 \text{ g} \times 1 \frac{\text{cal}}{\text{g K}} \times 393 \text{ K} = 3.93 \times 10^4 \text{ cal}$$

$$m_{H_2O} \lambda_e \approx 10^2 \text{ g} \times 540 \frac{\text{cal}}{\text{g}} = 5.40 \times 10^4 \text{ cal}$$

$$m_{H_2O} c_{H_2O} = 10^2 \frac{\text{cal}}{\text{K}}$$

$$m_o C_p = 5.2 \times 10^3 \frac{\text{cal}}{\text{K}}$$

$$\Rightarrow T_f = 305 \text{ K}$$

Supponiamo che $m_{H_2O} = 1 \text{ kg}$ (1 litro d'acqua)

$$\Rightarrow m_{H_2O} c_{H_2O} T_e = 3.93 \times 10^5 \text{ cal}$$

$$m_{H_2O} \lambda_e = 5.4 \times 10^5 \text{ cal} \Rightarrow T_f = 396 \text{ K} !$$

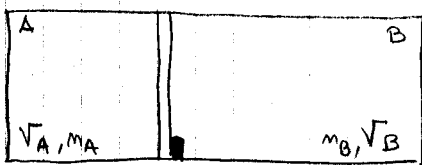
$$m_{H_2O} c_{H_2O} = 10^3 \text{ cal/K}$$

$T_f > T_e \Rightarrow$ ri-evapora! \Rightarrow la stanza non ce la fa a condensare il tutto.

Esercizio su gas perfetto

1

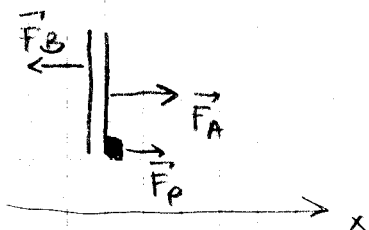
T_0 cost.



1) F esercitata dal pisto

2) dopo pisto, $V_{A,f}$ e $V_{B,f}$, P_f
(T resta cost.)

$$1) \quad P_A = \frac{m_A R T_0}{V_A} \quad ; \quad P_B = \frac{m_B R T_0}{V_B}$$

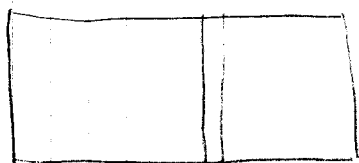


$$\vec{F}_A + \vec{F}_B + \vec{F}_p = 0$$

$$F_p = P_B S - P_A S$$

$$= \left(\frac{m_B}{V_B} - \frac{m_A}{V_A} \right) R T_0 S$$

2)



$$P_{A,f} = P_{B,f} = P_f$$

$$V_{A,f} + V_{B,f} = V_A + V_B$$

$$\frac{P_{A,f} V_{A,f}}{R T_0} = m_A$$

$$\Rightarrow \quad V_{A,f} = \frac{R T_0 m_A}{P_f} \quad ; \quad V_{B,f} = \frac{R T_0 m_B}{P_f}$$

$$\frac{P_{B,f} V_{B,f}}{R T_0} = m_B$$

$$\frac{R T_0}{P_f} (m_A + m_B) = V_A + V_B$$

$$\Rightarrow \quad P_f = R T_0 \frac{m_A + m_B}{V_A + V_B}$$

$$V_{A,f} = m_A \frac{V_A + V_B}{m_A + m_B}$$

$$V_{B,f} = m_B \frac{V_A + V_B}{m_A + m_B}$$